



TECHNICAL FACT SHEET– TUNGSTEN

At a Glance

- ❖ Hard, steel-gray to tin-white solid.
- ❖ Highest melting point among metals and is a good conductor of electricity.
- ❖ Typically used in welding, oil-drilling, electrical, and aerospace industries.
- ❖ Low solubility in water and high sorption (soil/water distribution) coefficients at low to neutral pH levels.
- ❖ Questions are being raised about tungsten's environmental stability.
- ❖ Exposure may cause eye and skin irritation, cough, nausea, diffuse interstitial pulmonary fibrosis, and changes in blood.
- ❖ No federal drinking water standard established.
- ❖ Exposure limits set by the National Institute for Occupational Safety and Health (NIOSH) and the American Conference of Governmental Industrial Hygienists (ACGIH).
- ❖ Treatment methods for tungsten in environmental media currently under development. Methods under investigation involve ice-electrodes, electrokinetic soil remediation, and chemical recovery/soil washing.

Introduction

This fact sheet, developed by the U. S. Environmental Protection Agency (EPA) Federal Facilities Restoration and Reuse Office (FFRRO), provides a brief summary for tungsten, including: physical and chemical properties; environmental and health impacts; existing federal and state guidelines; detection and treatment methods; and additional sources of information.

Tungsten was originally considered a metal that remains stable in soil and did not dissolve easily in water. However, it is now a growing concern to EPA and the U. S. Department of Defense (DoD) because recent research indicates that tungsten may not be as stable as was indicated in earlier studies. Furthermore, varying soil properties such as pH may cause tungsten to dissolve and leach into the underlying aquifer (ATSDR 2005). Currently, little information is available about the fate and transport of tungsten in the environment and its effects on human health. Research about tungsten is ongoing and includes health effects and risks, degradation processes, and an inventory of its use in the defense industry as a substitute for lead-based munitions. This fact sheet provides basic information on tungsten to site managers and other field personnel who may be faced with tungsten contamination at cleanup sites.

What is tungsten?

- ❖ Tungsten (also known as Wolfram and represented by the letter W in the periodic table) is a naturally occurring element that exists in the form of minerals or other compounds but typically not as a pure metal (ATSDR 2005; NIOSH 2007).
- ❖ Wolframite ($[\text{FeMn}]\text{WO}_4$) and Scheelite (CaWO_4) are two common minerals that contain tungsten (Koutsospyros et al. 2006).
- ❖ Based on its purity, the color of tungsten may range from white for the pure metal to steel-gray for the metal with impurities. It is commercially available in a powdered or solid form (ATSDR 2005; NIOSH 2007).
- ❖ The melting point of tungsten is the highest among metals and it resists corrosion. It is a good conductor of electricity and acts as a catalyst in chemical reactions (ATSDR 2005; Koutsospyros et al. 2006; Massachusetts DEP 2006).
- ❖ Tungsten powder is highly flammable and may ignite instantly on contact with air. Tungsten also may cause fire or explosion on contact with oxidants (ATSDR 2005; NIOSH 2007).
- ❖ DoD has used tungsten as a replacement for lead in bullets and other ammunition since 1999 (Massachusetts National Guard 2006).

What is tungsten (continued)?

- ❖ Tungsten ore is used primarily to produce tungsten carbide and tungsten alloys, which are used in many general welding and metal-cutting applications, in making drilling equipment for oil wells, and in operations within the aerospace industry. Tungsten metal is also used to produce lamp filaments, X-ray tubes, dyes, and paints for fabrics (ATSDR 2005; Koutsospyros et al 2006).
- ❖ Under the U. S. Army's Green Bullet program, nearly 88 million tungsten bullets were produced; 33 million remain unfired and available for use at training ranges across the country. Currently, the Army Environmental Center is looking at training ranges to evaluate the presence of various chemicals and other potential contaminants, including tungsten (Inside EPA 2007).

Exhibit 1: Physical and Chemical Properties of Tungsten
(HSDB 2009; NIEHS 2003; NIOSH 2007)

Property	Value
CAS Number	7440-33-7
Physical Description (physical state at room temperature)	Hard, steel-gray to tin-white solid
Molecular weight (g/mol)	183. 9
Water solubility (g/L at 25°C)	Insoluble at pH less than 6. 5
Boiling point (°C)	5,927
Melting point (°C)	3,410
Vapor pressure at 25°C (mm Hg)	0
Specific gravity	19. 3

Notes: g/mol – grams per mole; g/L – grams per liter; °C – degrees Celsius; mm Hg – millimeters of mercury.

What are the environmental impacts of tungsten?

- ❖ Tungsten is a common contaminant at industrial sites that use the metal and at DoD sites involved in the manufacture, storage, and use of tungsten-based ammunition. It is also found in detectable amounts in municipal solid waste and landfill leachate because of its use in common household products such as light bulbs (Koutsospyros et al. 2006).
- ❖ Tungsten particles may be present in air as a result of mining, weathering of rocks, or industrial applications that involve tungsten. These particles may settle on soil, water, or other surfaces and can be deposited through rain or other forms of precipitation (ATSDR 2005; MA DEP 2006).
- ❖ Tungsten is considered a “lithophilic” element, based on its strong soil binding capacity and its insolubility in water. Sorption coefficients increase at lower pH values, indicating lower mobility of tungsten (Koutsospyros et al. 2006).
- ❖ Increased acidification and oxygen depletion of soils from dissolution of tungsten powder have been shown to trigger changes in the soil microbial community, causing an increase in fungal biomass and a decrease in the bacterial component (Strigul et al. 2005).
- ❖ Studies indicate that an elevated pH in soil at a site may increase the solubility of tungsten by decreasing its sorption coefficient, which may cause it to leach more readily into the ground water table (ATSDR 2005; ASTSWMO 2008).
- ❖ Tungsten has been detected at six National Priorities List (NPL) sites (ATSDR 2005).
- ❖ In 2006, the assumed stability of tungsten in the environment was questioned when tungsten was detected in ground water and above baseline levels in soil at a small arms range at Massachusetts Military Reservation (MMR) where tungsten nylon bullets were being used. The use of these bullets was then suspended (ATSDR 2005; EPA 2010b; Massachusetts National Guard 2006).

What are the health effects of tungsten?

- ❖ Toxicological information on tungsten and its compounds is limited (Koutsospyros et al. 2006).
- ❖ Occupational exposure is considered the most common scenario for human exposure to tungsten and its compounds. Inhalation, ingestion, and dermal and eye contact are the possible exposure pathways (ATSDR 2005).

What are the health effects of tungsten? (continued)

- ❖ Occupational inhalation exposure to tungsten is known to affect the eyes, skin, respiratory system, and blood (ATSDR 2005). Tungsten may cause irritation to eyes, skin and throat; diffuse interstitial pulmonary fibrosis; loss of appetite; nausea; cough; and changes in the blood (NIOSH 2007).
- ❖ Tungsten was included as part of EPA's 2008 Integrated Risk Information System (IRIS) agenda. Toxicity is currently being assessed (EPA 2010a).
- ❖ Studies on rats have shown that oral exposure to tungsten caused post-implantation deaths and developmental abnormalities in the musculoskeletal system. Exposure of pregnant rats to sodium tungstate resulted in fetal death (NIEHS 2003).
- ❖ Tungsten has not been classified for carcinogenic effects by the Department of Health and Human Services (DHHS), the International Agency for Research on Cancer (IARC), or EPA (ATSDR 2005).

Are there any federal and state guidelines and health standards for tungsten?

- ❖ A federal drinking water standard has not been established for tungsten.
- ❖ The National Institute for Occupational Safety and Health (NIOSH) and the American Council of Government Industrial Hygienists (ACGIH) have established a recommended exposure limit (REL) of 5 milligrams per cubic meter (mg/m^3) as the time-weighted average (TWA) over a 10-hour work exposure and 10 mg/m^3 as the 15-minute, short-term exposure limit (STEL) for airborne exposure to tungsten (ATSDR 2005; ASTSWMO 2008).
- ❖ The Occupational Safety and Health Administration (OSHA) recommends an exposure limit of 5 mg/m^3 to insoluble compounds of tungsten and a 1 mg/m^3 limit of exposure to soluble compounds in construction and shipyard industries (ATSDR 2005).
- ❖ Massachusetts has established action levels of 1 to 2 parts per million (ppm) in soil and 15 ppm in groundwater (ASTSWMO 2008).

What detection and site characterization methods are available for tungsten?

- ❖ NIOSH Method 7074 – flame atomic absorption spectroscopy has a detection limit of 0.1 mg/m^3 for insoluble forms of tungsten and 0.05 mg/m^3 for soluble forms of tungsten in air (HSDB 2009; NIOSH 2007).
- ❖ Other NIOSH methods known to be used for tungsten are Methods 7300 and 7301, involving inductively coupled argon plasma, atomic emission spectroscopy, each with a detection limit of 0.005 mg/m^3 (NIOSH 2003a, b). Special sample treatment may be required for some tungsten compounds.
- ❖ ID213 – inductively coupled plasma atomic emission spectroscopy (ICP-AES) has a detection limit of 0.34 mg/m^3 for tungsten in air (NIOSH 2007; OSHA 2007).
- ❖ Tungsten in soil and water can be measured using the ICP-AES, ICP-mass spectrometry (ICP-MS), and ultraviolet/visible spectroscopy (UV/VIS) methods (ATSDR 2005).
- ❖ Tungsten is not currently included on the list of recoverable metals using SW-846 Method 3051. However, the digestion method has modified to enhance tungsten recovery from soils (Griggs et al. 2009).

What technologies are being used to treat tungsten?

- ❖ Treatment technologies to address tungsten contamination in environmental media are currently under development.
- ❖ According to preliminary studies conducted by various research groups, potential treatment methods involve chemical recovery/soil washing and phytoremediation (Lehr 2004; Warminsky and Larson 2004).
- ❖ "Ice electrode" is an innovative technology being evaluated for liquid media. This technology is based on the conventional electroplating technique and uses an electrode coated with a thin layer of ice. Tungsten ions adhere to the ice-coated electrode and can be removed by melting the ice (Lehr 2004).
- ❖ Electrokinetic soil remediation is an emerging in situ technology for removal of tungsten from low-permeability soils in the presence of copper and lead. A direct current is applied to contaminated soils using electrodes inserted into the ground (Braida et al. 2007).

Where can I find more information about tungsten?

- ❖ Agency for Toxic Substances and Disease Registry (ATSDR). 2005. Toxicological Profile for Tungsten. www.atsdr.cdc.gov/toxprofiles/tp186.pdf
- ❖ Association of State and Territorial Solid Waste Management Officials (ASTSWMO). 2008. Tungsten Issues Paper. www.astswmo.org/Files/Policies_and_Publications/Federal_Facilities/TUNGSTEN_FINAL_120208.pdf
- ❖ Braidă, W., C. Christodoulatos, A. Ogundipe, D. Dermatas, and G. O'Connor. 2007. Electrokinetic treatment of firing ranges containing tungsten-contaminated soils. *Journal of Hazardous Materials*. Volume 149. Pages 562 to 567.
- ❖ Griggs C., S. Larson, C. Nestler, and M. Thompson. 2009. Coupling of oxygen and pH requirements for effective microwave-assisted digestion of soils for tungsten analysis. *Land Contamination & Reclamation*. Volume 17. Pages 121 to 128.
- ❖ Hazardous Substances Data Bank (HSDB): Tungsten Compounds. 2009. <http://toxnet.nlm.nih.gov/cgi-bin/sis/htmlgen?HSDB>
- ❖ Inside EPA. 2007. Defense Environment Alert – Emerging Contaminants: Army risk study on tungsten could be first step toward regulation. January 23, 2007.
- ❖ Koutsospyros, A., W. Braidă, C. Christodoulatos, D. Dermatas, and N. Strigul. 2006. A review of tungsten: From environmental obscurity to scrutiny. *Journal of Hazardous Materials*. Volume 136. Pages 1 to 19.
- ❖ Lehr, Jay H. 2004. Wiley's Remediation Technologies Handbook: Major Contaminant Chemicals and Chemical Groups (ISBN 0-471-45599-7). Pages 282, 567, & 568.
- ❖ Massachusetts Department of Environmental Protection (DEP). 2006. Fact Sheet: Tungsten and Tungsten Compounds.
- ❖ Massachusetts National Guard. 2006. Massachusetts National Guard Temporarily Suspends Use of Tungsten-Nylon Ammunition at Camp Edwards, Massachusetts – News Release.
- ❖ National Institute of Environmental Health Sciences (NIEHS). 2003. Tungsten and Selected Tungsten Compounds – Review of Toxicological Literature.
- ❖ National Institute for Occupational Safety and Health (NIOSH). 2003a. Elements by ICP (Nitric/Perchloric Acid Ashing) – Method 7300. NIOSH Manual of Analytical Methods (NMAM), Fourth edition. www.cdc.gov/niosh/pdfs/97-162-f.pdf
- ❖ NIOSH. 2003b. Elements by ICP (Aqua Regia Ashing) – Method 7301. NIOSH Manual of Analytical Methods (NMAM), Fourth Edition. www.cdc.gov/niosh/docs/2003-154/chaps.html
- ❖ NIOSH. 2007. NIOSH Pocket Guide to Chemical Hazards: Tungsten. www.cdc.gov/niosh/docs/2005-149/pdfs/2005-149.pdf
- ❖ Occupational Safety and Health Administration (OSHA). 2007. Tungsten and Cobalt in Workplace Atmospheres (ICP Analysis).
- ❖ Strigul, N., A. Koutsospyros, P. Arienti, C. Christodoulatos, D. Dermatas, and W. Braidă. 2005. Effects of tungsten on environmental systems. *Chemosphere*. Volume 61. Pages 248 to 258.
- ❖ U. S. Environmental Protection Agency (EPA). 2010a. IRIS Agenda and Literature Searches. <http://cfpub.epa.gov/ncea/cfm/recordisplay.cfm?deid=187215>
- ❖ EPA. 2010b. Region 1. Massachusetts Military Reservation. www.epa.gov/region01/mmr/.
- ❖ Warminsky, Michael F. and Dr. Steven Larson. Recovery and Recycling of Tungsten and Lead from Small Arms Firing Range Soils. 2004. Presented at the Annual Conference on Soil, Sediment, and Water – University of Massachusetts Fall Conference (October 19-21). www.umasssoils.com/abstracts2004/Tuesday/trainin-granges.htm.

Contact Information

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